|  |
| --- |
| Chemical Investigations |
| Effect of Concentration on Reaction Speed |
| Teacher/Technician Guide |

A picture containing indoor, tub, bathtub

Description automatically generatedA picture containing room, drawing

Description automatically generated

Effect of Concentration on Reaction Speed

*UNIT 1 PPA 2*

**INTRODUCTION**

**The aim of this experiment is to find out how changing the concentration of sulphuric acid changes the speed that it reacts with magnesium.**

When a piece of magnesium reacts with sulphuric acid bubbles of gas are formed. The piece of magnesium gets smaller and smaller until it 'disappears'. If we time how long it takes to 'disappear' we can get some idea of the reaction speed. The **longer** the time the **slower** the reaction speed.

We will do three experiments but only one factor - **the concentration of the sulphuric acid** - will **change** in each one. To change the concentration of the sulphuric acid we will add water to it - this dilutes the acid and makes it less concentrated. Other factors must be kept the **same**. This means that the pieces of magnesium ribbon we use must be the **same** size and the experiments must be carried out at the **same** temperature. Each experiment will be done at room temperature

**You will need**

|  |  |
| --- | --- |
| 100 cm3 glass beaker | 2 mol l-1 sulphuric acid |
| 50 cm3 glass beaker | 2 cm long pieces of clean magnesium ribbon |
| 10 cm3 syringe | 20 cm3 syringe |
| timer |  |

**Safety**

Wear goggles.

Do not sniff at the reaction mixture – an irritating acid mist is formed.

Keep away from flames.

A picture containing icon

Description automatically generated**Procedure (what you do)**

1. Add dilute sulphuric acid to one of the 100 cm3 glass beakers until it is half full. The acid has a concentration of 2 mol l-1. Record this concentration by writing it down in the table
2. Diagram

   Description automatically generatedUsing the 20 cm3 syringe measure out 20 cm3 of the acid into the 50 cm3 glass beaker.

A picture containing icon

Description automatically generated

1. Add a piece of magnesium ribbon to the acid and at the same time start the timer.

A picture containing diagram

Description automatically generated

1. When the magnesium has just 'disappeared' stop the timer and record the time to the nearest second.
2. Wash out the small beaker and dry it.
3. Fill the other 100 cm3 beaker half full with water.
4. Using the 10 cm3 syringe measure out 10 cm3 of water into the small dry beaker. Then using the 20 cm3 syringe measure out 10 cm3 of the acid into the same beaker. This makes the concentration of the acid 1 mol l-1. Record this new concentration in the table on your 'assessment' sheet.
5. Add a piece of magnesium ribbon to the diluted acid and measure and record the time it takes for the magnesium to just 'disappear'.
6. Repeat the experiment using 15 cm3 of water and 5 cm3 of acid.
7. Measure and record the time it takes for the magnesium to just 'disappear'.
8. The acid concentration this time is 0.5 mol l-1. Record this concentration in the table

**Results sheet**

*What was the aim of the experiment?*

**Procedure**

*How did you change the concentration of the sulphuric acid?*

*How did you get some idea of how quickly the magnesium reacted with the dilute sulphuric*

*acid?*

*Complete the following table:*

*Graphical user interface, text, application

Description automatically generated with medium confidence*

**Conclusion**

*What did you find out from this experiment?*