**SSERC Risk Assessment** (revised version November 2009)

(based on HSE ‘5 steps to risk assessment’)

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| Activity assessed | Belousov-Zhabotinskii reaction |
| *Date of assessment* | 8/20/2018 |
| *Date of review (****Step 5****)* |  |
| *School* |  |
| *Department* |  |

| Step 1 | Step 2 | Step 3 | | Step 4 | | |
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| *List Significant hazards here:* | *Who might be harmed and how?* | *What are you already doing?* | *What further action is needed?* | *Action by whom?* | *Action by when?* | *Done* |
| Potassium bromate is (and sodium bromated probably is) a carcinogen as well as being toxic if ingested and oxidizing.  Malonic acid and iron II sulphate are harmful and an irritant  6M H2S04 is corrosive  1,10 phenanthroline is toxic if ingested  Reaction mixture gives off bromine.  Reaction mixture is corrosive, due to sulphuric acid, and carcinogenic, due to the bromate. | Teacher/technician preparing solutions by inhalation or splashes.  Teacher/technician preparing solutions by inhalation or splashes.  Teacher/technician preparing/using solution by splashes.  Teacher/technician preparing solutions by inhalation or splashes.  Teacher/technician or bystanders preparing solutions by inhalation  Teacher/technician carrying out reaction or clearing away by splashes. | Wear goggles (BS EN166 3) and gloves. Avoid raising dust. Keep away from combustible materials  Wear eye protection. Avoid raising dust.  Wear goggles (BS EN166 3) or a face shield for diluting concentrated acid and follow appropriate guidance. Wear goggles (BS EN166 3) when using solution.  Wear goggles (BS EN166 3) and gloves. Avoid raising dust.  The amount of bromine released is small. Work in a well-ventilated laboratory or in a fume cupboard.  Wear goggles (BS EN 166 3) and gloves. |  |  |  |  |

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| **Description of activity:**  To carry out the demonstration place about 5 cm3 of the bromate in a beaker and add 1 cm3 each of the bromide, malonic acid and sulphuric acid..  The mixture will go yellow because of bromine production. Swirl the beaker until the yellow colour completely disappears. Then add 1 or 2 cm3of E (the ferroin indicator solution), and swirl to mix. Pour the mixture into a Petri dish.  Initially red in colour, let the solution stand. After a while, tiny blue spots will start to appear. Each spot will slowly expand, eventually producing a series of concentric rings. The reaction may continue for half an hour, or even longer. |

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| **Additional comments:**  This reaction looks particularly good when the pertri dish with the mixture is placed on an overhead projector and the image projected on a screen. |